

Regiocomplementary Cycloaddition Reactions of Boryl- and Silylbenzynes with 1,3-Dipoles: Selective Synthesis of Benzo-Fused Azole Derivatives

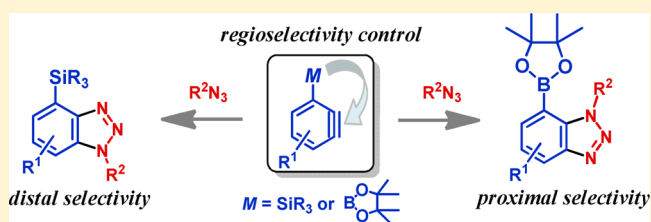
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Supporting Information

ABSTRACT: Benzo-fused nitrogen-containing heterocycles are abundant in biologically active compounds. One of the most important methods for preparing such heterocycles is the (3 + 2) cycloaddition reaction of benzynes with 1,3-dipolar compounds. However, the reactions of unsymmetrically substituted benzynes generally show low selectivity and hence yield mixtures of two regioisomers. In this paper, we describe the synthesis of both regioisomers of multisubstituted benzo-fused azole derivatives such as benzotriazoles, 1*H*-indazoles, and benzo[*d*]isoxazoles through the regiocomplementary (3 + 2) cycloaddition reactions of 3-boryl- and 3-silylbenzynes with 1,3-dipoles. The improved generation of 3-borylbenzynes from new precursors was one of the most important results of this work, which produced the successful (3 + 2) cycloaddition reactions with exclusive and proximal selectivities. On the other hand, similar reactions of 3-silylbenzynes selectively afforded distal cycloadducts. Analysis of the reaction pathways of these amazing regioselectivities by density functional theory calculations revealed that the (3 + 2) cycloadditions of borylbenzynes are controlled by the electrostatic effect of the boryl group, while those of silylbenzynes are controlled mainly by the steric effect of the bulky silyl groups that produced electrostatically unfavorable adducts via anomalous transition states.



INTRODUCTION

Benzo-fused azole derivatives are one of the most important classes of nitrogen-containing heterocycles.¹ For example, they have been reported to have a wide range of biological activities such as anticancer,² antidepressant,³ antidiabetic,⁴ antimicrobial,⁵ antifungal,⁶ antitubercular,⁷ and antipolytropic activities.⁸ These compounds have traditionally been synthesized using a linear, stepwise transformation of substituted benzene derivatives.⁹ However, these methods usually involve many steps and are less effective for the preparation of drug candidate libraries because each compound must be synthesized from a different starting material.

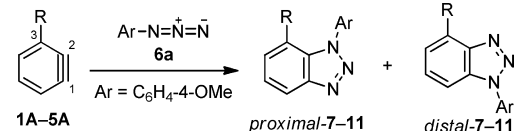
Cycloadditions of benzynes with 4*π*-components have attracted much attention recently. Among them, the (3 + 2) cycloaddition of benzynes with 1,3-dipole compounds has served as an effective alternative route to benzo-fused azole derivatives.¹⁰ In 2007, Yamamoto et al. reported a synthesis of 1*H*-indazole^{10b} using the (3 + 2) cycloaddition of benzynes, generated from 2-(trimethylsilyl)phenyl triflate with a fluoride source, to diazo compounds. Subsequently, a large number of benzo-fused azole derivatives such as 1*H*-indazole,^{10d,h,j} 2*H*-indazole,^{10k} benzotriazole,^{10c,f,g} and benzo[*d*]isoxazole^{10a,e,i,l} were synthesized using benzyne reactions. These transformations are valuable for convergent diversity-oriented

syntheses of heterocyclic compounds and allow effective preparation of drug candidate libraries.

However, like most benzyne reactions,¹¹ the (3 + 2) cycloaddition of unsymmetrically substituted benzynes with 1,3-dipoles is not always regioselective. For example, the reaction of 3-methylbenzynes **1A** with 4-methoxyphenyl azide **6a** yields a 1:1 mixture of two regioisomers of benzotriazole (*proximal*- and *distal*-**7Aa**) (Table 1, entry 1),^{10g} and a similar reaction of 3-fluorobenzynes **2A** also exhibited poor regioselectivity (entry 2).^{10g} The reaction of 3-methoxybenzynes **3A** is one of the special cases in which the (3 + 2) cycloaddition with **6a** exclusively gave *distal*-**9Aa** (entry 3),^{10g} and similar regioselective (3 + 2) reactions were also reported using 3-alkoxybenzynes.^{10b,c,e,g,h,j} In addition, Suzuki et al. discovered another kind of the regioselective (3 + 2) cycloaddition reaction using (3-trialkylsilyl)benzynes with nitron.^{10a} However, since then only a few groups have focused on the regiocontrol of the (3 + 2) cycloadditions of benzynes possessing a substituent other than an alkoxy group,^{10e,12,13} although so many new 1,3-dipolar cycloadditions have been reported over the last decade.¹⁰ Moreover, 3-silylbenzynes have

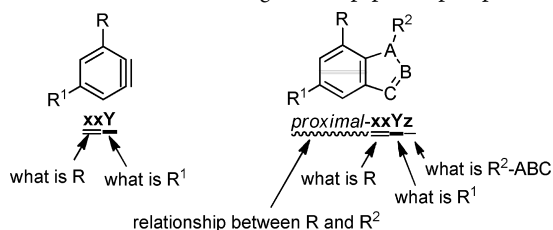
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Table 1. Regioselectivities of (3 + 2) Cycloadditions of 3-Substituted Benzyne 1A–5A with 4-Methoxyphenyl Azide 6a


entry	R	1–5 ^a	7–11 ^a	proximal:distal	ref
1	Me	1A	7Aa	1:1	10g
2	F	2A	8Aa	1:1	10g
3	OMe	3A	9Aa	1:>50	10g
4	Bpin ^b	4A	10Aa	>50:1	this study
5	SiMe ₃	5A	11Aa	1:5.7	this study

^aGeneral rules for the numbering on this paper. ^bBpin: pinacolboryl



been generated under harsh reaction conditions using *n*-BuLi.^{10a,e,12}

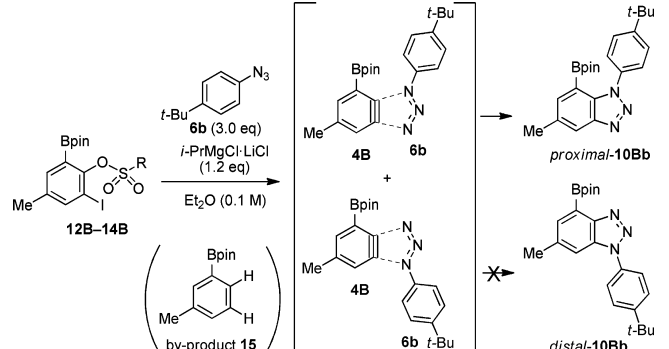
We recently reported that 3-boryl- and 3-silylbenzyne undergo Diels–Alder reactions with substituted furans and pyrroles to provide cycloadducts with high distal regioselectivity.¹⁴ We applied these benzyne to the (3 + 2) cycloadditions with various 1,3-dipoles and discovered the following. The cycloaddition reactions of 3-borylbenzyne 4 selectively gave proximal benzo-fused azoles 10, whereas similar reactions of 3-silylbenzyne 5 gave distal products 11. These opposite regioselectivities were also generally obtained with a range of 1,3-dipoles (typical examples in Table 1, entries 4 and 5; for more examples, see Table 3). We applied these protocols to the regioselective synthesis of the biologically interesting triazole analogue of hippadine, vorozole, and its regioisomer. To reveal the origin of the opposite regioselectivities of the (3 + 2)

cycloadditions of 4 and those of 5, we performed theoretical analysis of the reaction pathways by density functional theory (DFT) calculations. Although a part of this work was reported in our recent brief account,¹⁵ this paper describes the full details of this work. In particular, it provides new information about the reasons for the improvement of the borylbenzyne precursor and its development process, the experimental procedures and spectroscopic data of the benzyne precursors and products, the preparation of an aza analog of hippadine, and a full discussion about regioselectivities. Our previously published account only summarized the results without discussing them in detail, and in this paper we discuss the details of each reaction, pointing out many important features of the reactions that were not mentioned in the previous account.

RESULTS

Development of Improved Preparation Methods of 3-Boryl- and 3-Silylbenzyne. The most critical problems addressed in this project are the development of better precursors for the boryl- and silylbenzyne (4 and 5) and the optimization of the conditions for benzyne generation followed by (3 + 2) cycloaddition, because our preliminary trials, in which the generation methods for 4 and 5 (developed in our previous Diels–Alder reactions^{14b}) were applied to the (3 + 2) cycloaddition with 4-(*tert*-butyl)phenyl azide 6b, were not very successful. For example, *i*-PrMgCl·LiCl (1.2 equiv) was added to an Et₂O solution of 6-boryl-2-iodo-4-methylphenyl triflate 12B and 6b at –78 °C, and the reaction mixture was stirred at –78 °C for 30 min to generate 4B and promote the cycloaddition. We observed the formation of a cycloaddition product, *proximal*-10Bb as a single regioisomer; however, its yield was as low as 15%, and a significant amount of byproduct 15 (29%) along with the recovery of 12B (26%) were observed (Table 2, entry 1). We suspected that 15 was produced by the competitive hydride reduction¹⁶ of 4B by another *i*-PrMgCl molecule, which would leave behind a substantial amount of 12B.

Therefore, to improve the yield of *proximal*-10Bb, we attempted to suppress the formation of 15 using two approaches. First, we used Grignard reagents without β-

Table 2. Efficient Preparation of 3-Borylbenzyne 4B for (3 + 2) Cycloaddition to 4-(*tert*-Butyl)phenyl Azide 6b


entry	R	12B–14B	conditions	yield (%) of <i>proximal</i> -10Bb ^a
1	CF ₃	12B	–78 °C, 30 min	15 ^b
2	C ₆ H ₄ -4-Cl	13B	–78 °C, 30 min then 0 °C, 30 min	48
3	C ₆ H ₂ -2,4,5-Cl ₃	14B	–78 °C, 30 min then 0 °C, 30 min	67 ^c

^aThe yield of *proximal*-10Bb was determined by ¹H NMR; *distal*-10Bb was not detected in a crude product of each entry. ^bThe formation of 15 (29% yield) and the recovery of 12B (26% yield) were also observed. ^cIsolated yield of *proximal*-10Bb.

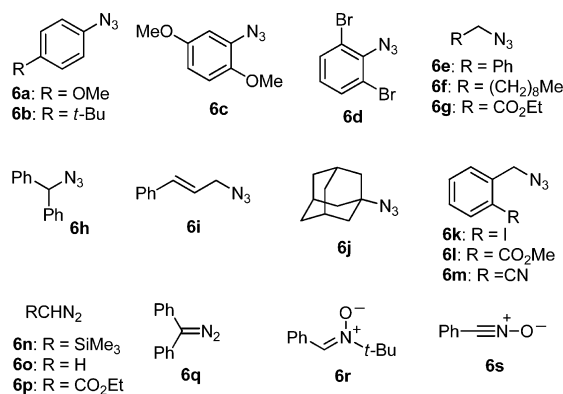


Figure 1. Structure of 1,3-dipoles 6a–6s.

hydrogen atoms, such as MeMgCl, BnMgCl, and *t*-Bu₂CHMgCl, instead of *i*-PrMgCl; however, there was no improvement (see Table S1 in Supporting Information). We next examined leaving groups other than –OSO₂CF₃. Knochel et al. reported a very important method for preparing functionalized benzyne from 2-iodophenyl 4-chlorobenzenesulfonate using the reaction with *i*-PrMgCl.¹⁷ The use of corresponding 4-chlorobenzenesulfonate **13B** significantly improved the yield of *proximal*-**10Bb** by up to 48% (Table 2, entry 2). After extensive studies on the benzyne precursors (see Tables S2 and S3 in Supporting Information), we finally found that 2,4,5-trichlorobenzenesulfonate **14B**^{18,19} was the best precursor of **4B** and that its (3 + 2) cycloaddition with **6b** afforded *proximal*-**10Bb** in 67% yield (Table 2, entry 3). It was proved that the iodine–magnesium exchange reaction of **14B** proceeded rapidly at –78 °C; however, the subsequent formation of **4B** did not proceed at the same temperature when **14B** was used as a precursor (see Tables S2 and S3 in Supporting Information). **4B** was gradually generated while the reaction mixture was warmed to 0 °C and immediately reacted with **6b**. Thus, the leaving ability of the sulfonyloxy group of the benzyne precursor affected the rate of benzyne generation, and tuning this leaving ability was the key to suppressing the undesired hydride reduction and allowing the successful reaction of **4** with **6b**. The regioisomer *distal*-**10Bb** was not detected in any of the cases examined (Table 2), indicating the first successful proximal-selective (3 + 2) cycloaddition of 3-substituted benzyne with azides.

Previously, we prepared silylbenzyne **5** through the reaction of 2-bromo-6-silylphenyl triflates and *n*-BuLi and applied them in regioselective Diels–Alder reactions.^{14a} In this study, we found the generation of **5** from 2,6-bis(trimethylsilyl)phenyl triflates **13** using Bu₄NF^{14c,20} was more suitable for the (3 + 2) cycloaddition of **5** because of the milder reaction conditions required. In fact, the reaction of the in situ generated 5-methyl-3-(trimethylsilyl)benzyne **5B** with **6b** selectively yielded *distal*-**11Bb** (*distal:proximal* = 10:1) (Table 3, entry 1).¹² It is worth noting that the regioselectivity of **5B** was opposite to that of the above-mentioned borylbenzyne **4B**. This cannot be explained by the electrostatic effects of boron and silicon, which are both less electronegative than carbon (Allred-Rochow electronegativities: B, 2.0; Si, 1.7; C, 2.5).²¹

Scope and Limitation of (3 + 2) Cycloadditions of Boryl- and Silylbenzyne with Various 1,3-Dipoles.¹⁵ We next investigated the (3 + 2) cycloadditions of 3-borylbenzyne (**4A** and **4B**) and 3-silylbenzyne **5A–5D** with diverse 1,3-dipoles including azides **6a–6m**, diazo compounds **6n–6q**,

nitrene **6r**, and nitrile oxide **6s** (Figure 1 and Table 3). The (3 + 2) cycloaddition of **4** generally afforded *proximal*-**10**, while that of **5** selectively afforded *distal*-**11**. Many examples of opposite selectivities are presented in entries 1–9 of Table 3. On the other hand, nitrene **6r**¹³ underwent cycloaddition with both **4** and **5** to give *distal* 2,3-dihydrobenzo[*d*]isoxazoles (**10Br** and **11Br**) with excellent selectivity (entry 10; similar reactions of **5** with **6r** were previously reported by Suzuki^{10a} and Danishefsky^{10e} under harsh conditions). This is the only exception in which both **4** and **5** showed the same *distal* selectivity.

The (3 + 2) cycloaddition of **4** with Me₃SiCHN₂ **6n** exclusively gave a desilylated indazole, *proximal*-**10Bn** (entry 9), while similar reactions of **5** with diazo compounds (**6p** and **6q**) gave *distal* 1*H*-indazoles (**11Bp** and **11Bq**) (entries 18 and 19). Silylbenzyne **5B** and benzonitrile oxide **6s** were simultaneously generated from their precursors (**16B** and *N*-hydroxybenzimidoyl chloride, respectively) using Bu₄NF in a single flask and exclusively gave *distal* benzo[*d*]isoxazole **11Bs** (entry 20).

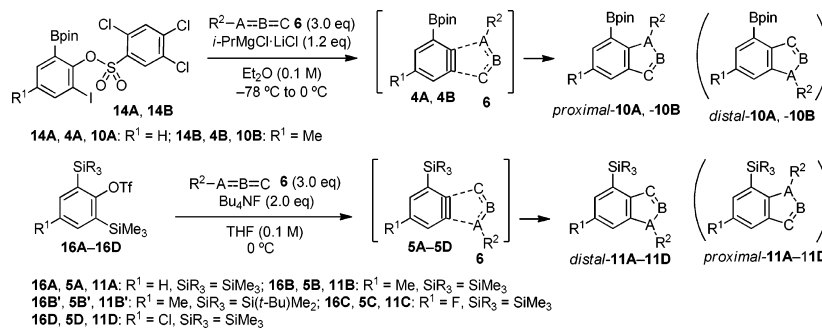
The following features are also worth noting: (i) Functional groups such as olefin (**6i**), ester (**6g** and **6l**), bromo (**6d**), iodo (**6k**), and cyano (**6m**) groups were tolerated under the reaction conditions when either *i*-PrMgCl·LiCl (for generating **4**) or a fluoride source (for generating **5**) was used. (ii) The reactions between **4** and most of the azides, including the bulky diphenylmethyl azide **6h**, resulted in complete proximal selectivity, and the reaction with the extremely bulky adamantyl azide **6j** also exhibited slight proximal preference (*proximal*-**10Bj**:*distal*-**10Bj** = 1.1:1) (Table 3, entry 8). (iii) The use of **5B'**, which has the more bulky *tert*-butyldimethylsilyl group, resulted in better *distal* selectivity (*distal*-**11Be'**:*proximal*-**11Be'** = 10:1) than did the use of 3-(trimethylsilyl)benzyne **5B** (entry 4). Similar results were also observed in the reaction with **6n** (entry 9). The *distal* selectivities also increased with increasing steric bulk of the alkyl chains in the azides (**6e**, **6h**, and **6j**; compare entries 4, 7, and 8). (iv) The substituent R¹ (H, Me, F, Cl) at the C5 position of **4** and **5** had little effect on the selectivity of the reaction with **6e** (entries 3, 4, 16, and 17).

Synthetic Applications of the (3 + 2) Cycloadducts.

The proximal-selective cycloaddition of 3-borylbenzyne **4** was applied to the synthesis of triazole analogues (**18A** and **18B**) of a testicular function inhibitor, hippadine **19**²² (Scheme 1a). When a suitably functionalized azide derivative **6t** was used, the (3 + 2) cycloaddition of borylbenzyne, generated from the corresponding precursors (**14A** and **14B**), gave benzotriazoles (*proximal*-**10At** and *proximal*-**10Bt**) in 69% and 67% isolated yields, respectively, with exclusive proximal selectivities. Their intramolecular Suzuki–Miyaura coupling afforded **17A** and **17B**, which were oxidized by MnO₂ to give **18A** and **18B** in 53% and 61% isolated yields, respectively, in two steps.

The formal synthesis of vorozole **21**⁹ (Scheme 1b) was performed using the *distal*-selective cycloaddition of 3-silylbenzyne **5**.¹⁵ First, 3-(*tert*-butyldimethylsilyl)benzyne, generated from **16E'**²³ by Bu₄NF, reacted with trimethylsilylmethyl azide **6u** to give *distal*-**11Eu'** (84% isolated yield) along with *proximal*-**11Eu'** (6% yield). Desilylation of *distal*-**11Eu'**, followed by deacetalization, afforded **20** (90% overall yield), a key intermediate for the synthesis of biologically active compounds such as aromatase inhibitor **21**⁹ and a PI3K inhibitor.²⁴ Moreover, we synthesized a regioisomer **23** of vorozole **21** via the proximal-selective (3 + 2) cycloaddition of a 3-borylbenzyne, generated from **14C**, with **6u** (Scheme 1c).

Table 3. Complementary Regioselective (3 + 2) Cycloaddition of Borylbenzynes 4 and Silylbenzynes 5

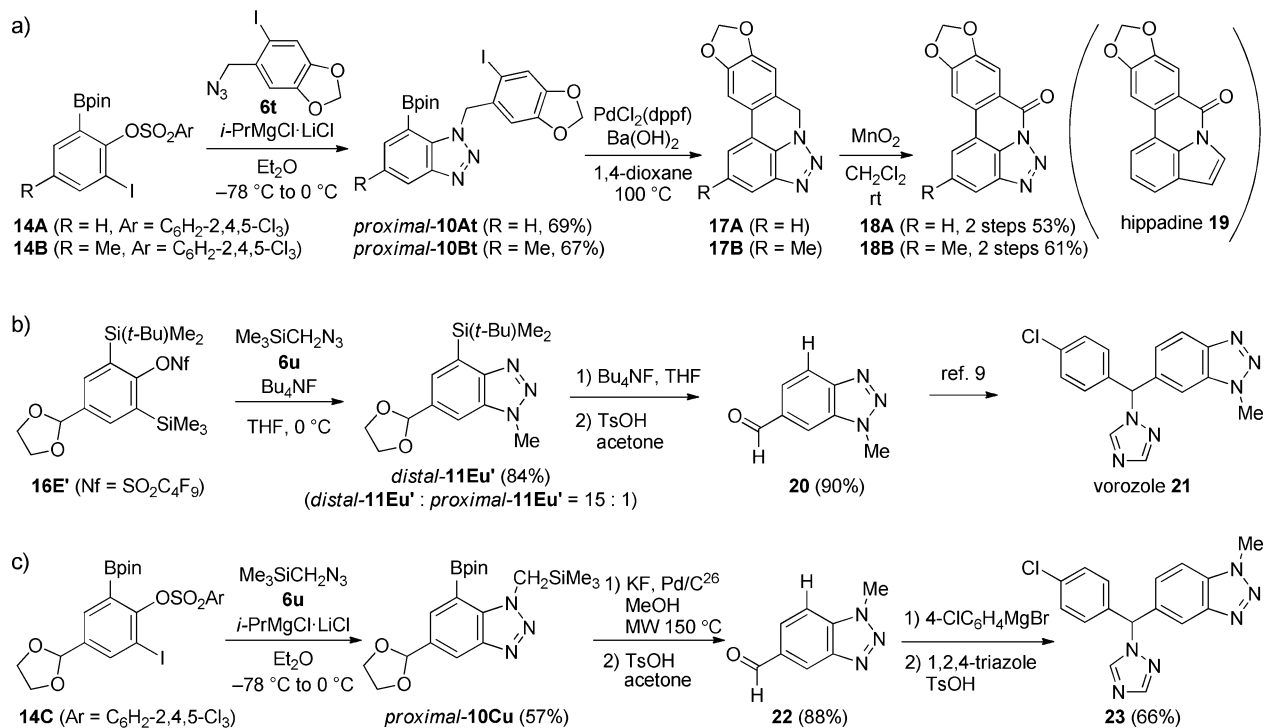


Entry	10	11	Entry	10	11
1			6		
	proximal-10Bb ^a (67% from 14B, 6b)	distal-11Bb (80% ^b from 16B, 6b) [10 : 1] ^c		proximal-10Bg ^a (74% from 14B, 6g)	distal-11Bg (78% ^b from 16B, 6g) [3.3 : 1] ^{c,e}
2			7		
	proximal-10Bc ^a (65% from 14B, 6c)	distal-11Bc (74% ^b from 16B, 6c) [14 : 1] ^c		proximal-10Bh ^a (65% from 14B, 6h)	distal-11Bh (65% ^b from 16B, 6h) [5.0 : 1] ^c
3			8		
	proximal-10Ae ^a (89% from 14A, 6e)	distal-11Ae (52% ^b from 16A, 6e) [3.6 : 1] ^c		proximal-10Bj (total 63% from 14B, 6j) [1.1 : 1] ^c	distal-11Bj (57% ^b from 16B, 6j) [>50 : 1] ^c
4			9		
	proximal-10Be ^a (85% from 14B, 6e)	distal-11Be [73% ^b (77% ^{b,d} from 16B, 6e) [3.3 : 1], ^c [10 : 1] ^{c,d}		proximal-10Bn ^a (54% from 12B, 6n)	distal-11Bn [98% ^b (99% ^{b,d} from 16B, 6n) [1.4 : 1], ^c [3.3 : 1] ^{c,d}
5			10		
	proximal-10Bf ^a (87% from 14B, 6f)	distal-11Bf (73% ^b from 16B, 6f) [4.0 : 1] ^c		distal-10Br ^f (64% from 12B, 6r)	distal-11Br (94% from 16B, 6r) [20 : 1] ^c

Other examples:

Entry 11	Entry 12	Entry 13	Entry 14	Entry 15
proximal-10Bd ^a (67% from 14B, 6d)	proximal-10Bi ^a (80% from 14B, 6i)	proximal-10Bk ^a (71% from 14B, 6k)	proximal-10Bl ^a (74% from 14B, 6l)	proximal-10Bm ^a (78% from 14B, 6m)
Entry 16	Entry 17	Entry 18	Entry 19	Entry 20
distal-11C (84% ^b from 16C, 6e) [2.8 : 1] ^{c,e}	distal-11D (44% ^b from 16D, 6e) [3.0 : 1] ^{c,e}	distal-11P (81% ^b from 16B, 6p) [12 : 1] ^c	distal-11Bq (85% ^b from 16B, 6q) [21 : 1] ^c	distal-11Bs (73% from 16B, 6s) [>50 : 1] ^c

^aProximal regioisomer **10** was exclusively observed in ¹H NMR spectra of the crude product. ^bTotal isolated yield of distal and proximal isomers **11**. ^cThe ratio of distal to proximal products, determined using ¹H NMR spectra of the crude product, is shown in brackets. ^dResults for 3-(*tert*-butyldimethylsilyl)benzynes **5B'** [$\text{SiR}_3 = \text{Si}(t\text{-Bu})\text{Me}_2$], which was used instead of 3-(trimethylsilyl)benzynes **5B** ($\text{SiR}_3 = \text{SiMe}_3$) and provided a mixture of the corresponding *distal*- and *proximal*-**11B'** [$\text{SiR}_3 = \text{Si}(t\text{-Bu})\text{Me}_2$]. The ratio of *distal*- to *proximal*-**11B'** [$\text{SiR}_3 = \text{Si}(t\text{-Bu})\text{Me}_2$] is shown in brackets. ^eOnly the corresponding desilylated product was isolated for *proximal*-**11**. ^fDistal regioisomer **10Br** was exclusively observed in ¹H NMR spectra of the crude product. Ad = 1-adamantyl.

Scheme 1. Application of [3 + 2] Cycloaddition to the Syntheses of Biologically Active Compounds and Their Aza Analogues^a

^a(a) Synthesis of hippadine analogues (18A and 18B). (b) Synthesis of vorozole 21. (c) Synthesis of a regioisomer 23 of vorozole 21.

These regioselective syntheses are valuable because the methylation of norvorozole produced a 1:1:1 mixture of 21 and its two regioisomers.²⁵ Our method involves fewer steps and allows the convergent synthesis of benzo-fused azoles from benzyne precursors and 1,3-dipoles. Thus, it should be useful for the combinatorial synthesis of diverse derivatives of biologically active compounds.

DISCUSSION

Herein, we discuss why the (3 + 2) cycloadditions of borylbenzynes 4 gave proximal products and those of silylbenzynes 5 gave distal products using the results of density functional theory (DFT) calculations.²⁷

First, we checked the internal angles of the geometry-optimized 3-boryl- (4A and 4B) and 3-silylbenzynes (5A–5D, and 5B') to evaluate the contribution of distortion energy,¹² and we also analyzed the charge distributions of these benzynes by performing natural population analysis in the isolated system (without solvent effect and coordination of metals) (Table 4).²⁸ Two important features of the benzynes were found: (i) The internal angles and the charge distribution of all these benzynes are almost exactly the same, and the substituent R¹ at the C5-position of benzyne has little effect. (ii) C2s are more electrophilic than C1s because of the electrostatic effect^{10a,14,15} of the boryl and silyl groups and also because of the benzyne distortion.^{12,29}

We next analyzed the experimental regioselectivities of the (3 + 2) cycloaddition reactions between two benzynes (4B and 5B) and three 1,3-dipoles (6e, 6o, and 6r) (for natural charges of these 1,3-dipoles, see Figure 2) to find whether they were consistent with their calculated charge distributions and benzyne distortions. The experimental regioselectivities of the reactions between borylbenzyne 4B and the 1,3-dipoles (6e, 6o, and 6r) are consistent with the calculated results, while those

Table 4. Internal Angle and Natural Charge of 3-Borylbenzynes 4 and 3-Silylbenzynes 5

Bpin	R ¹	4	internal angle (deg)		natural charge	
			C1	C2	C1	C2
	H	4A	122	133	-0.042	0.118
	Me	4B	122	132	-0.034	0.111

SiR ₃	R ¹	5	internal angle (deg)		natural charge	
			C1	C2	C1	C2
SiMe ₃	H	5A	121	135	-0.037	0.081
SiMe ₃	Me	5B	121	134	-0.030	0.076
SiMe ₃	F	5C	121	135	-0.027	0.080
SiMe ₃	Cl	5D	121	135	-0.027	0.092
SiMe ₃	Me	5B'	121	134	-0.031	0.072
SiH ₃	Me	5B''	123	132	-0.006	0.067

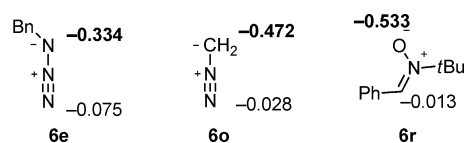
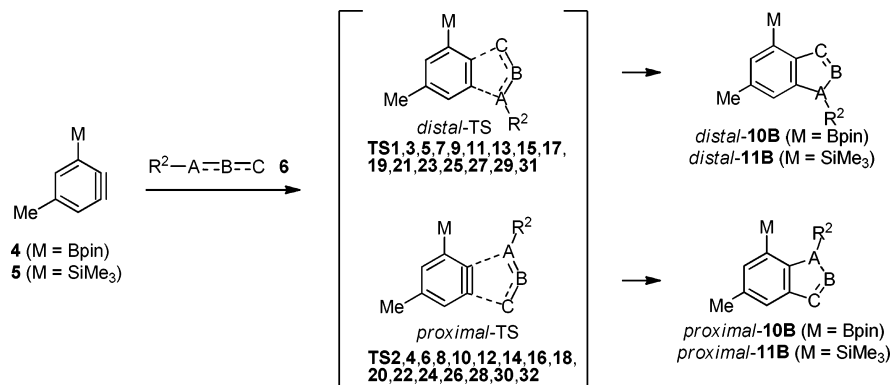


Figure 2. Natural charges of 1,3-dipolar compounds (6e, 6o, and 6r).

between silylbenzyne 5B and the 1,3-dipoles (6e and 6o) are not (Table 3, entries 4, 9, and 10). These contrasting results clearly indicate that the regioselectivities are accounted for by neither the electrostatic effect nor the benzyne distortion.

Table 5. Theoretical and Experimental Ratios in the (3 + 2) Cycloaddition Reactions of Borylbenzynes 4 and Silylbenzynes 5



entry	M	4, 5	6	TS1–TS32	theoretical results		exptl ratio of distal to proximal from Table 3
					$\Delta\Delta H^\ddagger$ or $\Delta\Delta G^\ddagger$ (kcal/mol) ^a	ratio of distal to proximal ³⁰	
1	Bin	4B	6e	TS1, TS2	+2.58 ^b	1:>50	1:>50
2	SiMe ₃	5B	6e	TS3, TS4	-1.11 ^c	6.5:1	3.3:1
3	SiH ₃	5B''	6e	TS5, TS6	-0.14 ^c	1.3:1	
4	Bpin	4B	6f	TS7, TS8	+2.22 ^b	1:43	1:>50
5	Bpin	4B	6g	TS9, TS10	+3.94 ^b	1:>50	1:>50
6	Bpin	4B	6h	TS11, TS12	+1.96 ^b	1:28	1:>50
7	Bpin	4B	6i	TS13, TS14	+2.36 ^b	1:>50	1:>50
8	Bpin	4B	6j	TS15, TS16	+0.63 ^b	1:2.9	1:1.1
9	Bpin	4B	6k	TS17, TS18	+2.78 ^b	1:>50	1:>50
10	Bpin	4B	6l	TS19, TS20	+3.02 ^b	1:>50	1:>50
11	Bpin	4B	6m	TS21, TS22	+3.60 ^b	1:>50	1:>50
12	SiMe ₃	5B	6c	TS23, TS24	-1.45 ^c	11:1	14:1
13	SiMe ₃	5B	6j	TS25, TS26	-2.40 ^c	>50:1	>50:1
14	SiMe ₃	5B	6p	TS27, TS28	-1.66 ^c	20:1	>50:1
15	SiMe ₃	5B	6q	TS29, TS30	-3.08 ^c	>50:1	>50:1
16	SiMe ₃	5B	6s	TS31, TS32	-2.73 ^c	>50:1	>50:1

^aEnergy difference between *distal*-TS and *proximal*-TS. Positive $\Delta\Delta H^\ddagger$ or $\Delta\Delta G^\ddagger$ indicates that *distal*-TS is higher than *proximal*-TS, while negative $\Delta\Delta H^\ddagger$ or $\Delta\Delta G^\ddagger$ represents the opposite. ^bActivation energy difference is shown as a $\Delta\Delta H^\ddagger$ value. ^cActivation energy difference is shown as a $\Delta\Delta G^\ddagger$ value.

Therefore, we performed a theoretical analysis of the reaction pathways, including the transition states, to quantitatively rationalize the origin of the selectivities.

The transition states (TS1–TS4) of the reactions between benzyl azide 6e and two benzynes (3-borylbenzynes 4B and 3-silylbenzynes 5B) were obtained as typical cases (Table 5, entries 1 and 2). TS2, which leads to *proximal*-10Be, has an energy 2.58 kcal/mol lower than that of TS1, which leads to *distal*-10Be. This energy difference corresponds to 1:>50 proximal selectivity, which is in good agreement with our experimental result (*distal*-10Be:*proximal*-10Be = 1:>50; Table 3, entry 4 and Table 5, entry 1). On the other hand, TS3, which leads to *distal*-11Be, has an energy 1.1 kcal/mol lower than that of TS4, which leads to *proximal*-11Be. This difference corresponds to 6.5:1 distal selectivity, which is in reasonable agreement with our experimental result (*distal*-11Be:*proximal*-11Be = 3.3:1; Table 3, entry 4 and Table 5, entry 2).³⁰

We also performed similar calculations for thirteen other reactions between 4B (or 5B) and various 1,3-dipolar compounds 6. All of these results are in good agreement with the experimental data (Table 5, entries 4–16). These facts suggest that the calculated transition states are reliable and that their structures can provide valuable information that can be used to identify the origin of the regioselectivities.³⁰

As a reference, the transition state TS33 of the reaction of benzyne 24A and 6e was also calculated, in which the bond distance (2.40 Å) between the internal nitrogen of 6e and the benzyne carbon is shorter than that (2.77 Å) between the terminal nitrogen and the benzyne carbon (Figure 3a). This is despite the fact that there should be some steric interaction between the benzylic methylene moiety of 6e and the benzyne hydrogen. The corresponding bond distances of the TS34, derived from 5-methylbenzynes 24B, and 6e are close to those of TS33, which implies that a methyl group at the C5-position of benzyne hardly affects the structure of the transition state. The lengths of the two bonds forming in TS2, which is more stable than TS1 because the electrostatic combination of the reacting atoms is well matched (Figure 3b), are also very similar to those of TS34. These facts imply that there is little steric interaction between the boryl group and the benzyl substituent and that the exclusive formation of *proximal*-10Be is dominated by the electrostatic interactions between the two reactants and also by the benzyne distortion, both of which are caused by the boryl group.¹⁵

In the case of silylbenzynes 5B, the distance between the internal nitrogen and the benzyne C2 is much longer in TS4 (2.51 Å) than in TS34 (2.41 Å) because of the strong steric interaction between the trimethylsilyl group and the benzyl group. This makes the electrostatically and distortionally

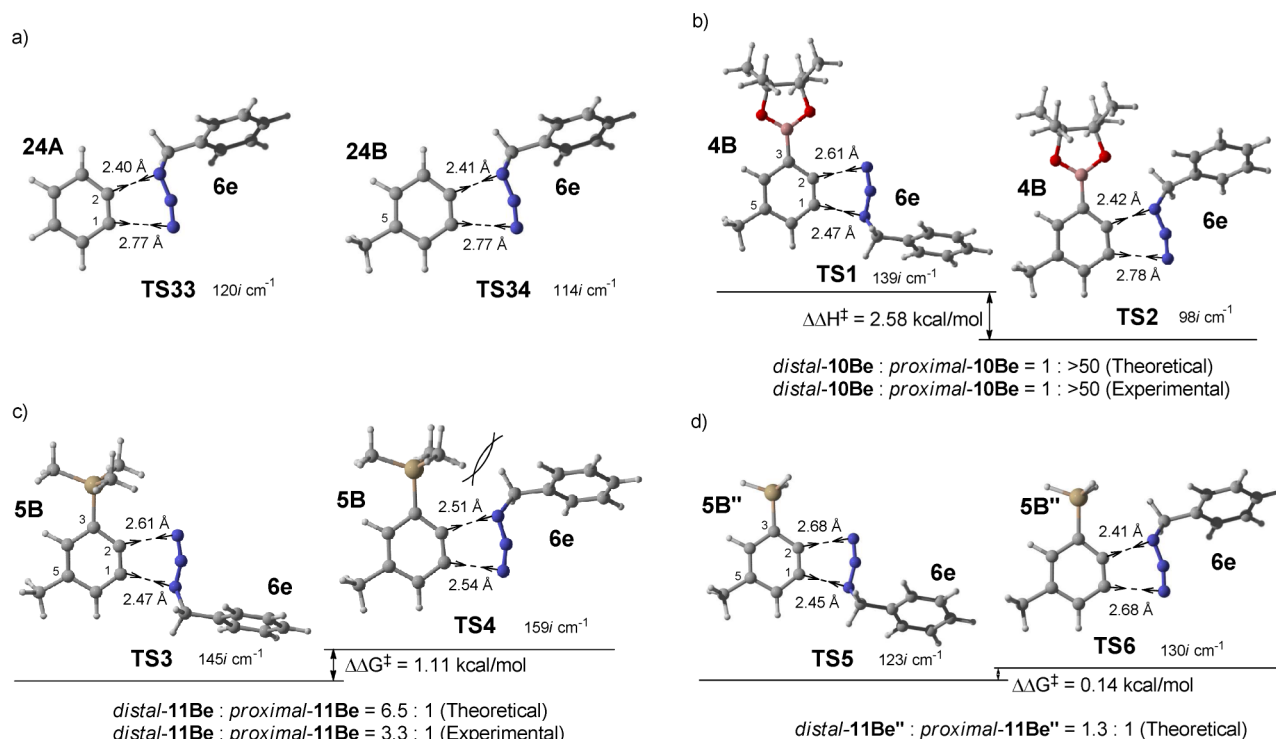


Figure 3. Transition states of (3 + 2) cycloaddition reactions of benzynes with 1,3-dipoles.^{15,32} (a) Transition states of the reactions between benzyne 24A and 6e (TS33) and between 5-methylbenzyne 24B and 6e (TS34). (b) Most probable transition states (TS1 and TS2) of the reaction between 3-boryl-5-methylbenzyne 4B and benzyl azide 6e giving *distal*-10Be and *proximal*-10Be, respectively. (c) Most probable transition states (TS3 and TS4) of the reaction between 5-methyl-3-(trimethylsilyl)benzyne 5B and 6e giving *distal*-11Be and *proximal*-11Be, respectively. (d) Most probable transition states (TS5 and TS6) of the reaction between 5-methyl-3-silylbenzyne 5B'' and 6e.

favorable TS4 less stable than the unfavorable TS3 (Figure 3c) and results in the preferential formation of *distal*-11Be. We also analyzed two transition states (TS5 and TS6) derived from a virtual benzyne 5B'' possessing a SiH₃ group (Figure 3d). The distance between the internal nitrogen and the benzyne C2 of TS6 (2.41 Å) is shorter than that of TS4 (2.51 Å) but similar to that of TS34 (2.41 Å). Although the benzyne distortion of 5B'' is quite similar to that of 5B (Table 4), the energy difference between TS5 and TS6 is smaller than that between TS3 and TS4 (Figure 3c and d). These results also prove that the steric bulkiness of the silyl groups has a significant adverse influence on the electrostatically favorable proximal transition state. This steric influence even overrides the attractive electrostatic interaction and the benzyne distortion.^{15,31}

CONCLUSIONS

In conclusion, we have achieved complementary regiocontrol of the (3 + 2) cycloaddition reaction of benzynes with 1,3-dipoles by the unique substituent effects of the boryl and silyl groups. Both regioisomers of benzo-fused azoles such as benzotriazole, 1*H*-indazole, and benzo[*d*]isoxazole can be prepared by choosing a boryl or silyl group as the benzyne substituent. In particular, the finding of an improved generation method of 3-borylbenzynes 4 from new precursors 14 was one of the most important new results of this work and led to the successful (3 + 2) cycloaddition reactions to demonstrate the first proximal-selective examples of the cycloaddition reaction of 3-substituted benzynes with 1,3-dipoles. We have clearly and quantitatively explained the regioselectivities of these reactions by analyzing the transition state structures obtained by DFT calculations. Namely, the (3 + 2) cycloaddition of borylbenzynes 4 is more

electrostatically controlled, while that of silylbenzynes 5 is more sterically dominated. Because the boryl and silyl groups of the cycloaddition products can be converted into carbon, nitrogen, and oxygen substituents as well as a hydrogen atom,³³ the use of 3-boryl- and 3-silylbenzynes would be the ideal solution to the longstanding problem of the regioselectivity of substituted benzynes. Application of this chemistry to the preparation of a wide range of biologically interesting compounds and further improvement of the selectivity are in progress in our laboratory.

EXPERIMENTAL SECTION

General. All reactions were carried out under an argon or nitrogen atmosphere in an oven-dried flask containing a stir-bar with a rubber septum or an inlet adapter with a three-way stopcock. 1-Azido-4-methoxybenzene 6a,³⁴ 1-azido-4-(*tert*-butyl)benzene 6b,³⁵ 2-azido-1,4-dimethoxybenzene 6c,^{10c} 2-azido-1,3-dibromobenzene 6d,³⁶ 1-azido-decane 6f,³⁷ ethyl 2-azidoacetate 6g,^{10c} (azidomethylene)dibenzene 6h,³⁸ (*E*)-(3-azidoprop-1-en-1-yl)benzene 6i,^{10c} (1*R*,3*R*,5*S*)-1-azido-damantane 6j,^{10c} 1-(azidomethyl)-2-iodobenzene 6k,^{10c} methyl 2-(azidomethyl)benzoate 6l,³⁹ 2-(azidomethyl)benzotrile 6m,⁴⁰ (diazomethylene)dibenzene 6q,⁴¹ *N*-hydroxybenzimidoyl chloride (for the in situ preparation of benzonitrile *N*-oxide 6s),⁴² 5-(azidomethyl)-6-iodobenzo[*d*]1,3-dioxole 6t,⁴³ 2-iodo-6-(4,4,5,5-tetramethyl-1,3-dioxolan-2-yl)phenyl trifluoromethanesulfonate 12A,^{14b} 2-iodo-4-methyl-6-(4,4,5,5-tetramethyl-1,3-dioxolan-2-yl)phenyl trifluoromethanesulfonate 12B,^{14b} 2,6-bis(trimethylsilyl)phenyl trifluoromethanesulfonate 16A,^{14c} 4-methyl-2,6-bis(trimethylsilyl)phenyl trifluoromethanesulfonate 16B,^{14c} 4-fluoro-2,6-bis(trimethylsilyl)phenyl trifluoromethanesulfonate 16C,^{14c} and 4-chloro-2,6-bis(trimethylsilyl)phenyl trifluoromethanesulfonate 16D^{14c} were synthesized according to the literature. Flash chromatography⁴⁴ was performed with silica gel 60 N, spherical neutral (40–50 μm). ¹H NMR and ¹³C NMR spectra were recorded on an instrument (¹H, 500 MHz; ¹³C, 125 MHz) with chemical shifts reported in ppm relative to

the residual deuterated solvent or the internal standard tetramethylsilane. The high-resolution mass spectra were recorded on an ESI or APCI TOF mass spectrometer. Yield refers to isolated yields of compounds greater than 95% purity as determined by ^1H NMR analysis. ^1H NMR and melting points (where applicable) of all known compounds were taken from references. The regioselectivities were determined by 500 MHz ^1H NMR spectra of crude reaction mixtures. Each regiochemistry of representative cycloaddition products (*proximal-10Bb*, *proximal-10Be*, *proximal-10Bf*, *proximal-10Bg*, *proximal-10Bh*, *proximal-10Bi*, *proximal-10Bj*, *proximal-10Bn*, *distal-10Br*, *distal-11Bc*, *distal-11Be*, *distal-11Bf*, *distal-11Bh*, *proximal-11Bh*, *distal-11Bj*, *distal-11Bn*, *distal-11Bp*, and *distal-11Bs*) was confirmed by NOESY experiment. The regiochemistries of all other cycloaddition products were predicted by the similarity of ^1H NMR data.

General Procedure I: Proximal-Selective (3 + 2) Cycloadditions of Borylbenzynes 4 (Tables 1–3). An oven-dried flask was charged with a borylbenzyne precursor **12**^{14b}–**14** (1.0 equiv) and a dipole **6** (3.0 equiv), capped with a rubber septum, then evacuated, and backfilled with argon. Anhydrous Et₂O (0.10 M) was added, and the mixture was cooled to $-78\text{ }^\circ\text{C}$. Then, a 1.0 or 1.3 M solution of *i*-PrMgCl·LiCl in THF (1.2 equiv) was slowly added to the reaction mixture over 5 min. After stirring at $-78\text{ }^\circ\text{C}$ for 30 min, the reaction mixture was warmed to $0\text{ }^\circ\text{C}$, stirred for an additional 30 min, and quenched with a saturated aqueous NH₄Cl solution. Then the reaction mixture was extracted with EtOAc, and the aqueous layer was extracted twice with EtOAc. The combined organic layers were washed with a saturated aqueous NaCl solution and dried over anhydrous Na₂SO₄. The solution was filtered through a glass filter and concentrated under reduced pressure. The crude product was purified by silica gel flash column chromatography to give boryl benzazole **10**.

1-(4-Methoxyphenyl)-7-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-1H-benzo[d]triazole (*proximal-10Aa*) (Table 1, entry 4). Following the general procedure I, a mixture of **14A**^{14b} (61 mg, 103 μmol), 1-azido-4-methoxybenzene **6a**³⁴ (46 mg, 0.31 mmol), and *i*-PrMgCl·LiCl (1.0 M in THF, 0.12 mL, 0.12 mmol) in Et₂O (1.1 mL, 0.10 M) was stirred for 30 min at $0\text{ }^\circ\text{C}$. The crude product was purified by column chromatography (hexane/EtOAc = 3:1) to provide the titled compound *proximal-10Aa* (17 mg, 46%) as a colorless solid. Mp $142\text{--}145\text{ }^\circ\text{C}$. ^1H NMR (500 MHz, CDCl₃) δ : 1.08 (12 H, s), 3.89 (3 H, s), 7.03 (2 H, td, $J = 3.0, 9.0\text{ Hz}$), 7.40 (1 H, dd, $J = 7.0, 8.0\text{ Hz}$), 7.41 (2 H, td, $J = 3.0, 9.0\text{ Hz}$), 7.92 (1 H, dd, $J = 1.0, 7.0\text{ Hz}$), 8.20 (1 H, dd, $J = 1.0, 8.0\text{ Hz}$). ^{13}C NMR (125 MHz, CDCl₃) δ : 24.5, 55.6, 83.8, 114.1, 122.7, 123.5, 128.1, 131.3, 136.3, 136.8, 145.3, 160.4. IR (CHCl₃): 1518 cm⁻¹. HRMS calcd for C₁₉H₂₂BN₃NaO₃ (M + Na)⁺ m/z : 374.1652, found 374.1661.

1-[4-(*tert*-Butyl)phenyl]-5-methyl-7-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-1H-benzo[d]triazole (*proximal-10Bb*) (Table 2, entry 1). An oven-dried pear-shaped flask was charged with **12B**^{14b} (44 mg, 90 μmol), capped with an inlet adapter with a three-way stopcock, then evacuated, and backfilled with argon. Anhydrous Et₂O (0.90 mL, 0.10 M) and 1-azido-4-(*tert*-butyl)benzene **6b**³¹ (62 mg, 0.35 mmol) were added via a syringe, and the mixture was cooled to $-78\text{ }^\circ\text{C}$. *i*-PrMgCl·LiCl (1.3 M in THF, 0.080 mL, 104 μmol) was slowly added into the reaction mixture over 5 min. After being stirred at $-78\text{ }^\circ\text{C}$ for 30 min, the reaction mixture was quenched with a saturated aqueous NH₄Cl solution. The mixture was extracted with EtOAc, and the aqueous layer was extracted twice with EtOAc. The combined organic layers were washed with a saturated aqueous NaCl solution and dried over anhydrous Na₂SO₄. The solution was filtered through a glass filter and concentrated under reduced pressure. The product yield and ratio were determined by ^1H NMR [*proximal-10Bb* (15%), **15**⁴⁵ (29%), **12B**^{14b} (26%), 1,4-dimethoxybenzene was used as an internal standard].

Table 2, entry 2. Following the general procedure I, a mixture of **13B** (25 mg, 48 μmol), 1-azido-4-(*tert*-butyl)benzene **6b**³⁵ (23 mg, 131 μmol), and *i*-PrMgCl·LiCl (1.3 M in THF, 60 μL , 78 μmol) in Et₂O (0.50 mL, 0.10 M) was stirred for 30 min at $0\text{ }^\circ\text{C}$. The product yield and ratio were determined by ^1H NMR [*proximal-10Bb* (48%), 1,4-dimethoxybenzene was used as an internal standard].

Table 2, entry 3 and Table 3, entry 1. Following the general procedure I, a mixture of **14B** (21 mg, 34 μmol), 1-azido-4-(*tert*-butyl)benzene **6b**³⁵ (20 mg, 113 μmol), and *i*-PrMgCl·LiCl (1.3 M in THF, 40 μL , 52 μmol) in Et₂O (0.40 mL, 0.10 M) was stirred for 30 min at $0\text{ }^\circ\text{C}$. The crude product was purified by column chromatography (hexane/EtOAc = 10:1) to provide the titled compound *proximal-10Bb* (8.9 mg, 67%) as a colorless solid, the regiochemistry of which was determined by NOESY spectrum. Mp $160\text{--}165\text{ }^\circ\text{C}$. ^1H NMR (500 MHz, CDCl₃) δ : 1.05 (12 H, s), 1.38 (9 H, s), 2.53 (3 H, s), 7.42 (2 H, d, $J = 8.5\text{ Hz}$), 7.53 (2 H, d, $J = 8.5\text{ Hz}$), 7.72 (1 H, d, $J = 1.0\text{ Hz}$), 7.95 (1 H, d, $J = 1.0\text{ Hz}$). ^{13}C NMR (125 MHz, CDCl₃) δ : 21.1, 24.5, 31.3, 31.3, 34.8, 83.8, 121.6, 125.9, 126.0, 133.4, 134.9, 135.8, 137.9, 146.1, 152.2. IR (CHCl₃): 2968 cm⁻¹. HRMS calcd for C₂₃H₃₁BN₃O₂ (M + H)⁺ m/z : 392.2509, found 392.2507.

1-(2,5-Dimethoxyphenyl)-5-methyl-7-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-1H-benzo[d]triazole (*proximal-10Bc*) (Table 3, entry 2). Following the general procedure I, a mixture of **14B** (31 mg, 52 μmol), 2-azido-1,4-dimethoxybenzene **6c**^{10c} (50 μL , 166 μmol), and *i*-PrMgCl·LiCl (1.3 M in THF, 50 μL , 65 μmol) in Et₂O (0.50 mL, 0.10 M) was stirred for 30 min at $0\text{ }^\circ\text{C}$. The crude product was purified by column chromatography (hexane/EtOAc = 10:1) to provide the titled compound *proximal-10Bc* (13 mg, 65%) as a colorless solid. Mp $116\text{--}120\text{ }^\circ\text{C}$. ^1H NMR (500 MHz, CDCl₃) δ : 1.07 (12 H, s), 2.52 (3 H, s), 3.58 (3 H, s), 3.79 (3 H, s), 6.93 (1 H, dd, $J = 1.0, 9.5\text{ Hz}$), 7.03–7.06 (2 H, m), 7.76 (1 H, d, $J = 1.0\text{ Hz}$), 7.95 (1 H, brs). ^{13}C NMR (125 MHz, CDCl₃) δ : 20.6, 24.8, 24.8, 37.3, 83.6, 88.9, 92.5, 114.6, 123.1, 123.7, 128.0, 128.3, 131.4, 134.6, 135.4, 1367.0, 138.4, 145.4. IR (CHCl₃): 1514 cm⁻¹. HRMS calcd for C₂₁H₂₇BN₃O₄ (M + H)⁺ m/z : 396.2095, found 396.2086.

1-Benzyl-7-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-1H-benzo[d]triazole (*proximal-10Ae*) (Table 3, entry 3). Following the general procedure I, a mixture of **14A** (64 mg, 0.11 mmol), (azidomethyl)benzene **6e** (41 μL , 0.32 mmol), and *i*-PrMgCl·LiCl (1.3 M in THF, 0.10 mL, 0.13 mmol) in Et₂O (1.1 mL) was stirred for 30 min at $0\text{ }^\circ\text{C}$. The crude product was purified by column chromatography (hexane/EtOAc = 10:1) to provide the titled compound *proximal-10Ae* (32 mg, 89%) as a colorless solid. Mp $89\text{--}92\text{ }^\circ\text{C}$. ^1H NMR (500 MHz, CDCl₃) δ : 1.25 (12 H, s), 6.40 (2 H, s), 7.01 (2 H, d, $J = 7.0\text{ Hz}$), 7.20–7.25 (3 H, m), 7.38 (1 H, dd, $J = 7.5\text{ Hz}, 8.0\text{ Hz}$), 8.04 (1 H, dd, $J = 1.0\text{ Hz}, 7.5\text{ Hz}$), 8.21 (1 H, dd, $J = 1.0\text{ Hz}, 8.0\text{ Hz}$). ^{13}C NMR (125 MHz, CDCl₃) δ : 24.6, 52.3, 84.4, 123.3, 123.5, 126.2, 127.3, 128.4, 135.8, 137.1, 137.3, 145.9. IR (CHCl₃): 3010 cm⁻¹. HRMS calcd for C₁₉H₂₃BN₃O₂ (M + H)⁺ m/z : 336.1883, found 336.1863.

1-Benzyl-5-methyl-7-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-1H-benzo[d]triazole (*proximal-10Be*) (Table 3, entry 4). Following the general procedure I, a mixture of **14B** (61 mg, 101 μmol), (azidomethyl)benzene **6e** (40 μL , 0.30 mmol), and *i*-PrMgCl·LiCl (1.3 M in THF, 90 μL , 117 μmol) in Et₂O (1.0 mL, 0.10 M) was stirred for 30 min at $0\text{ }^\circ\text{C}$. The crude product was purified by column chromatography (hexane/EtOAc = 4:1) to provide the titled compound *proximal-10Be* (30 mg, 85%) as a colorless solid, the regiochemistry of which was determined by NOESY spectrum. Mp $116\text{--}118\text{ }^\circ\text{C}$. ^1H NMR (500 MHz, CDCl₃) δ : 1.24 (12 H, s), 2.51 (3 H, s), 6.36 (2 H, s), 6.99 (2 H, d, $J = 7.0\text{ Hz}$), 7.17–7.24 (3 H, m), 7.86 (1 H, d, $J = 1.0\text{ Hz}$), 7.96 (1 H, brs). ^{13}C NMR (125 MHz, CDCl₃) δ : 21.1, 24.7, 52.3, 84.4, 122.5, 126.2, 127.2, 128.4, 133.2, 134.4, 137.2, 139.3, 146.7. IR (CHCl₃): 2982 cm⁻¹. HRMS calcd for C₂₀H₂₅BN₃O₂ (M + H)⁺ m/z : 350.2040, found 350.2037.

1-Decyl-5-methyl-7-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-1H-benzo[d]triazole (*proximal-10Bf*) (Table 3, entry 5). Following the general procedure I, a mixture of **14B** (56 mg, 93 μmol), 1-azidodecane **6f**³⁵ (60 mg, 0.33 mmol), and *i*-PrMgCl·LiCl (1.3 M in THF, 85 μL , 111 μmol) in Et₂O (0.90 mL, 0.10 M) was stirred for 30 min at $0\text{ }^\circ\text{C}$. The crude product was purified by column chromatography (hexane/EtOAc = 4:1) to provide the titled compound *proximal-10Bf* (32 mg, 87%) as a colorless solid, the regiochemistry of which was determined by NOESY spectrum. Mp $70\text{--}72\text{ }^\circ\text{C}$. ^1H NMR (500 MHz, CDCl₃) δ : 0.86 (3 H, t, $J = 7.5\text{ Hz}$),

117.7, 121.3, 123.2, 124.9, 127.8, 129.2, 131.4, 144.2, 149.9, 154.2, 155.4. IR (CHCl₃): 3008, 1715, 1039 cm⁻¹. HRMS calcd for C₁₄H₈N₃O₃ (M + H)⁺ *m/z*: 266.0566, found 266.0583.

Preparation of 2-Methyl-7H-[1,3]dioxolo[4,5-*f*]triazolo[4,5,1-*de*]phenanthridin-7-one (18B) from proximal-10Bt. An oven-dried round-bottom flask was charged with proximal-10Bt (50 mg, 0.096 mmol), PdCl₂(dppf)·CH₂Cl₂ (24 mg, 0.030 mmol), and Ba(OH)₂·8H₂O (91 mg, 0.30 mmol). The flask was capped with a rubber septum, then evacuated, and backfilled with argon. Anhydrous dioxane (0.50 mL) was added, and the mixture was stirred at 100 °C for 18 h. After the reaction, the mixture was filtered through a short pad of Celite, washed with CH₂Cl₂, and evaporated under reduced pressure to give crude 17B. Then, activated MnO₂ (80 mg, 0.96 mmol) was added into the crude 17B. The flask was capped with a rubber septum, then evacuated, and backfilled with argon. Anhydrous CH₂Cl₂ (3.3 mL) was added, and the mixture was stirred for 24 h at room temperature. After the reaction, the mixture was filtered through a short pad of Celite, washed with CH₂Cl₂, and evaporated under reduced pressure. The residue was washed with acetone to provide the titled compound 18B (16 mg, 61% from proximal-10Bt) as a pale yellow solid. Mp 256–258 °C. ¹H NMR (500 MHz, CDCl₃) δ: 2.69 (3 H, s), 6.22 (2 H, s), 7.60 (1 H, s), 7.93 (2 H, s), 7.98 (1 H, s), 8.06 (1 H, s). ¹³C NMR (125 MHz, CDCl₃) δ: 22.1, 102.2, 102.9, 109.5, 116.7, 120.7, 123.6, 124.4, 127.7, 130.9, 137.4, 144.9, 149.6, 153.8, 155.2. IR (CHCl₃): 3008, 1703, 1039 cm⁻¹. HRMS calcd for C₁₅H₁₀N₃O₃ (M + H)⁺ *m/z*: 280.0722, found 280.0727.

Formal Synthesis of Vorozole 21 (Scheme 1b). 4-(tert-Butyldimethylsilyl)-6-(1,3-dioxolan-2-yl)-1-methyl-1H-benzo[d]triazole (distal-11Eu'). Following the general procedure II, a mixture of 16E' (100 mg, 0.158 mmol), (azidomethyl)trimethylsilane 6u (62 mg, 0.48 mmol) (for safety reasons, commercially available 6u was used as an equivalent of the highly explosive methyl azide), and Bu₄NF (1.0 M in THF, 0.32 mL, 0.32 mmol) in THF (1.6 mL) was stirred for 30 min at 0 °C. The ratio of distal- to proximal-11Eu' (15:1) was determined using ¹H NMR spectra of the crude product. The crude product was purified by column chromatography (hexane/EtOAc = 3:1) to provide the titled compound distal-11Eu' (43 mg, 84%) as a colorless solid. Mp 92–94 °C. ¹H NMR (500 MHz, CDCl₃) δ: 0.52 (6 H, s), 0.93 (9 H, s), 4.09–4.20 (4 H, m), 4.28 (3 H, s), 6.00 (1 H, s), 7.54 (1 H, d, *J* = 8.0 Hz), 7.68 (1 H, d, *J* = 8.0 Hz). ¹³C NMR (125 MHz, CDCl₃) δ: -5.0, 17.2, 26.9, 34.1, 65.4, 103.5, 107.5, 129.5, 131.8, 133.2, 136.2, 151.0. IR (CHCl₃): 1602, 1408, 1328, 1236, 1120 cm⁻¹. HRMS calcd for C₁₆H₂₆N₃O₂Si (M + H)⁺ *m/z*: 320.1794, found 320.1801.

6-(1,3-Dioxolan-2-yl)-1-methyl-1H-benzo[d]triazole (24). Bu₄NF (1.0 M in THF, 0.35 mL, 350 μmol) was added into the THF solution (0.70 mL, 0.10 M) of distal-11E' (23 mg, 71 μmol) and stirred for 48 h at 70 °C. The reaction mixture was cooled to room temperature and evaporated under reduced pressure. The crude product was purified by column chromatography (hexane/EtOAc = 3:1) to provide the titled compound 24 (14 mg, 94%) as a colorless solid. Mp 97–98 °C. ¹H NMR (500 MHz, CDCl₃) δ: 4.08–4.18 (4 H, m), 4.31 (3 H, s), 5.98 (1 H, s), 7.49 (1 H, d, *J* = 8.0 Hz), 7.69 (1 H, s), 8.05 (1 H, d, *J* = 8.0 Hz). ¹³C NMR (125 MHz, CDCl₃) δ: 34.3, 65.4, 103.2, 107.1, 120.0, 122.6, 133.3, 137.7, 146.4. IR (CHCl₃): 1600, 1193 cm⁻¹. HRMS calcd for C₁₀H₁₂N₃O₂ (M + H)⁺ *m/z*: 206.0930, found 206.0910.

1-Methyl-1H-benzo[d]triazole-6-carbaldehyde (20). TsOH·H₂O (3.4 mg, 20 μmol) was added into the acetone solution (0.70 mL) of 24 (14 mg), and the mixture was stirred at room temperature for 3 h. After the reaction, a saturated aqueous NaHCO₃ was added into the reaction mixture, and the mixture was extracted with Et₂O. The aqueous layer was extracted twice with Et₂O. The combined organic layers were washed with a saturated aqueous NaCl solution. The organic layer was dried over anhydrous Na₂SO₄ and evaporated under reduced pressure. The crude product was purified by column chromatography (hexane/EtOAc = 3:1) to provide the titled compound 20 (10 mg, 96%) as a colorless solid. Mp 119–122 °C. ¹H NMR (500 MHz, CDCl₃) δ: 4.41 (3 H, s), 7.91 (1 H, d, *J* = 8.0 Hz), 8.11 (1 H, s), 8.19 (1 H, d, *J* = 8.0 Hz), 10.19 (1 H, s). ¹³C NMR

(125 MHz, CDCl₃) δ: 34.7, 112.4, 120.9, 124.0, 133.5, 135.1, 148.6, 191.3. IR (CHCl₃): 1703, 1616, 1602, 1236 cm⁻¹. HRMS calcd for C₈H₈N₃O (M + H)⁺ *m/z*: 162.0667, found 162.0692.

Synthesis of iso-Vorozole 23 (Scheme 1c). 5-(1,3-Dioxolan-2-yl)-7-(4,4,5,5-tetramethyl-1,3,2-dioxaborolan-2-yl)-1-((trimethylsilyl)methyl)-1H-benzo[d]triazole (proximal-10Cu). Following the general procedure I, a mixture of 14C (94 mg, 0.14 mmol), (azidomethyl)trimethylsilane 6u (0.21 mL, 1.41 mmol), and *i*-PrMgCl·LiCl (1.3 M in THF, 0.21 mL, 0.27 mmol) in Et₂O (1.4 mL) was stirred for 30 min at 0 °C. The crude product was purified by column chromatography (hexane/EtOAc = 5:2) to provide the titled compound proximal-10Cu (33 mg, 57%) as a colorless oil. ¹H NMR (500 MHz, CDCl₃) δ: 0.15 (9 H, s), 1.40 (12 H, s), 4.07–4.09 (2 H, m), 4.17–4.20 (2 H, m), 4.57 (2 H, s), 5.98 (1 H, s), 8.10 (1 H, d, *J* = 1.5 Hz), 8.24 (1 H, d, *J* = 1.5 Hz). ¹³C NMR (125 MHz, CDCl₃) δ: -2.1, 24.9, 40.7, 65.3, 84.5, 103.5, 121.3, 132.9, 135.2, 136.9, 145.3. IR (CHCl₃): 1587 cm⁻¹. HRMS calcd for C₁₉H₃₀BN₃NaO₄Si (M + Na)⁺ *m/z*: 426.1996, found 426.1991.

5-(1,3-Dioxolan-2-yl)-1-methyl-1H-benzo[d]triazole (25). The mixture of proximal-10Cu (27 mg, 67 μmol), 10% Pd/C (3.4 mg), and KF (12 mg, 0.20 mmol) was evacuated and backfilled with argon. MeOH (0.70 mL) was added into the reaction tube, and the mixture was heated using microwave for 3 h at 150 °C.²⁶ The reaction mixture was filtered through a short pad of Celite and evaporated under reduced pressure. CH₂Cl₂ was added into the mixture, filtered through a short pad of Celite, concentrated under reduced pressure to provide 25 as a colorless solid (13.5 mg), and used without further purification for the next reaction.

1-Methyl-1H-benzo[d]triazole-5-carbaldehyde (22). TsOH·H₂O (3.8 mg, 20 μmol) was added into an acetone solution (0.70 mL) of 25 (13.5 mg). The mixture was stirred at room temperature for 9 h. After the reaction, a saturated aqueous NaHCO₃ was added into the reaction mixture and extracted with EtOAc. The aqueous layer was extracted twice with EtOAc. The combined organic layers were washed with a saturated aqueous NaCl solution. The organic layer was dried over anhydrous Na₂SO₄ and evaporated under reduced pressure. The crude product was purified by flash column chromatography (hexane/EtOAc = 2:1) to provide the titled compound 22 (9.5 mg, 88% 2 steps) as a colorless solid. Mp 165–167 °C. ¹H NMR (500 MHz, CDCl₃) δ: 4.36 (3 H, s), 7.63 (1 H, d, *J* = 8.5 Hz), 8.09 (1 H, dd, *J* = 1.0, 8.5 Hz), 8.55 (1 H, brs), 10.13 (1 H, s). ¹³C NMR (125 MHz, CDCl₃) δ: 34.5, 110.1, 125.6, 126.1, 133.1, 136.5, 145.7, 191.1. IR (CHCl₃): 1701 cm⁻¹. HRMS calcd for C₈H₇N₃NaO (M + Na)⁺ *m/z*: 184.0487, found 184.0481.

(4-Chlorophenyl)(1-methyl-1H-benzo[d]triazol-5-yl)-methanol (26). An oven-dried flask was charged with 22 (46 mg, 0.29 mmol), capped with an inlet adapter with a three-way stopcock, then evacuated, and backfilled with argon. Et₂O (2.8 mL, 0.10 M) was added via a syringe into the reaction mixture and cooled to -78 °C. 4-Chlorophenylmagnesiumbromide (1.0 M in Et₂O, 0.57 mL, 0.57 mmol) was added into the solution and stirred at -78 °C for 1 h. The reaction mixture was quenched with a saturated aqueous NH₄Cl and extracted with EtOAc. The aqueous layer was extracted twice with EtOAc. The combined organic layers were washed with a saturated aqueous NaCl solution. The organic layer was dried over anhydrous Na₂SO₄ and evaporated under reduced pressure. The crude product was purified by column chromatography (hexane/EtOAc = 2:1) to provide the titled compound 26 (65 mg, 83%) as a colorless oil. ¹H NMR (500 MHz, CDCl₃) δ: 4.28 (3 H, s), 6.00 (1 H, s), 7.32 (4 H, m), 7.46 (2 H, s), 8.08 (1 H, s). ¹³C NMR (125 MHz, CDCl₃) δ: 34.3, 75.3, 109.5, 117.3, 126.6, 127.9, 128.7, 133.1, 133.5, 139.9, 142.0, 145.9. IR (CHCl₃): 3603, 1600, 1491, 1281, 1236 cm⁻¹. HRMS calcd for C₁₄H₁₃ClON₃ (M + H)⁺ *m/z*: 274.0747, found 274.0726.

5-((4-Chlorophenyl)(1H-1,2,4-triazol-1-yl)methyl)-1-methyl-1H-benzo[d]triazole (23). A mixture of 26 (65 mg, 0.23 mmol), 1,2,4-triazole (33 mg, 0.47 mmol), and TsOH·H₂O (13 mg, 71 μmol) in a pear-shaped flask was dissolved in toluene (5.0 mL). The flask was connected with a Dean–Stark apparatus and stirred at 150 °C for 24 h. The reaction mixture was cooled to room temperature and quenched with a saturated aqueous NaHCO₃. The reaction mixture was

compound **41** (1.20 g, 95%) as a colorless solid. Mp 103–104 °C. ¹H NMR (500 MHz, CDCl₃) δ: 0.46 (6 H, s), 1.07 (9 H, s), 4.05 (4 H, m), 5.65 (1 H, s), 7.86 (2 H, s). ¹³C NMR (125 MHz, CDCl₃) δ: 0.0, 18.9, 26.5, 65.3, 88.5, 101.5, 134.4, 138.5, 155.9. IR (CHCl₃): 2958, 2931, 2887, 2858, 1600, 1448, 1276, 1255 cm⁻¹. HRMS calcd for C₁₅H₂₃I₂O₃Si (M + H)⁺ *m/z*: 532.9528, found 532.9506.

2-(tert-Butyldimethylsilyl)-4-(1,3-dioxolan-2-yl)-6-iodophenol (42). An oven-dried round-bottom flask was charged with **41** (0.80 g, 1.51 mmol), capped with an inlet adapter with a three-way stopcock, then evacuated, and backfilled with nitrogen. THF (15 mL, 0.10 M) was added into the flask via a syringe, cooled to -78 °C, and stirred for 10 min. *n*-BuLi (1.65 M in hexane, 1.2 mL, 2.0 mmol) was added to the solution at -78 °C, warmed to room temperature, and stirred for 2 h. The reaction mixture was quenched with a saturated aqueous NH₄Cl and extracted with EtOAc. The aqueous layer was extracted twice with EtOAc. The combined organic layers were washed with a saturated aqueous NaCl solution. The organic layer was dried over anhydrous Na₂SO₄ and concentrated under reduced pressure. The crude product was purified by column chromatography (hexane/EtOAc = 10:1) to provide the titled compound **42** as a colorless solid (0.53 g, 88%). Mp 112–113 °C. ¹H NMR (500 MHz, CDCl₃) δ: 0.30 (6 H, s), 0.89 (9 H, s), 0.78 (9 H, s), 4.06 (4 H, m), 5.53 (1 H, s), 5.70 (1 H, s), 7.36 (1 H, d, *J* = 2.0 Hz), 7.80 (1 H, d, *J* = 2.0 Hz). ¹³C NMR (125 MHz, CDCl₃) δ: -4.8, 17.6, 27.0, 65.2, 87.2, 102.8, 123.5, 131.2, 135.7, 137.4, 159.4. IR (CHCl₃): 3489, 2955, 2928, 2886, 2856, 1589, 1560, 1413, 1359, 1248 cm⁻¹. HRMS calcd for C₁₅H₂₄IO₃Si (M + H)⁺ *m/z*: 407.0539, found 407.0551.

2-(tert-Butyldimethylsilyl)-4-(1,3-dioxolan-2-yl)-6-(trimethylsilyl)phenol (43). An oven-dried round-bottom flask was charged with **42** (63 mg, 0.15 mmol), capped with an inlet adapter with a three-way stopcock, then evacuated, and backfilled with nitrogen. CH₂Cl₂ (1.0 mL, 0.25 M) was added into the flask via a syringe, cooled to 0 °C, and stirred for 10 min. Triethylamine (26 μL, 0.19 mmol) and trimethylsilyl chloride (25 μL, 0.19 mmol) were added into the flask, and the mixture was stirred at room temperature for 1 h. The reaction mixture was evaporated under reduced pressure. Hexane was added to the residue, filtered through a pad of short pad of Celite and concentrated under reduced pressure to afford (5-(1,3-dioxolan-2-yl)-3-iodo-2-((trimethylsilyl)oxy)phenyl)(*tert*-butyl)-dimethylsilane. The crude product was dissolved in THF (1.0 mL, 0.25 M) and stirred at -78 °C for 10 min. *n*-BuLi (1.5 M in hexane, 0.12 mL, 0.19 mmol) was added to the mixture and stirred at room temperature for 1 h. The reaction mixture was quenched with a saturated aqueous NH₄Cl and evaporated under reduced pressure. The residue was extracted with hexane and water. The aqueous layer was extracted twice with hexane. The combined organic layers were washed with a saturated aqueous NaCl solution. The organic layer was dried over anhydrous Na₂SO₄ and solvent was removed under reduced pressure. The crude product was purified by column chromatography (hexane/EtOAc = 20:1) to provide the titled compound **43** (23 mg, 42%) as a colorless oil. ¹H NMR (500 MHz, CDCl₃) δ: 0.32 (9 H, s), 0.39 (6 H, s), 0.92 (9 H, s), 4.08 (4 H, m), 5.16 (1 H, s), 5.74 (1 H, s), 7.41 (1 H, d, *J* = 2.0 Hz), 7.50 (1 H, d, *J* = 2.0 Hz). ¹³C NMR (125 MHz, CDCl₃) δ: -4.6, -0.9, 17.6, 26.5, 65.2, 104.2, 120.3, 125.0, 128.6, 134.9, 136.3, 166.5. IR (CHCl₃): 3609, 2955, 1587, 1576, 1409, 1364, 1238 cm⁻¹. HRMS calcd for C₁₈H₃₃O₃Si₂ (M + H)⁺ *m/z*: 353.1968, found 353.1982.

2-(tert-Butyldimethylsilyl)-4-methyl-6-(trimethylsilyl)phenyl 1,1,2,2,3,3,4,4,4-Nonafluorobutane-1-sulfonate (16E'). Compound **43** (4.4 g, 12.5 mmol) and 18-crown-6 (3.3 g, 12.5 mmol) were dissolved in THF (125 mL, 0.10 M) and stirred at 0 °C for 10 min. NaH (0.75 g, 18.7 mmol) was added to the mixture and stirred at 0 °C for 30 min. Then, nonafluorobutanesulfonyl fluoride²³ (0.62 mL, 3.7 mmol) was added to the solution and stirred at 80 °C for 20 h. The reaction mixture was quenched with water and evaporated under reduced pressure. The residue was extracted with EtOAc and water, and the aqueous layer was extracted twice with hexane. The combined organic layers were washed with a saturated aqueous NaCl solution. The organic layer was dried over anhydrous Na₂SO₄, filtered with a glass filter, and concentrated under reduced pressure (this reaction did

not complete). THF (125 mL, 0.10 M), NaH (0.25 g, 6.2 mmol) and nonafluorobutanesulfonyl fluoride (1.2 mL, 7.1 mmol) were added again to obtain a full conversion. The mixture was stirred at 80 °C for 20 h. The reaction mixture was quenched with water and evaporated under reduced pressure. The residue was extracted with EtOAc and water, and the aqueous layer was extracted twice with hexane. The combined organic layers were washed with a saturated aqueous NaCl solution. The organic layer was dried over anhydrous Na₂SO₄, filtered with a glass filter and evaporated under reduced pressure. The crude product was purified by column chromatography (hexane/EtOAc = 5:1) to provide the titled compound **16E'** (6.1 g, 77%) as a colorless solid. Mp 67–69 °C. ¹H NMR (500 MHz, CDCl₃) δ: 0.35 (9 H, s), 0.40 (6 H, s), 0.79 (9 H, s), 4.05–4.12 (4 H, m), 5.84 (1 H, s), 7.67 (2 H, s). ¹³C NMR (125 MHz, CDCl₃) δ: -4.0, 0.5, 18.1, 26.9, 65.3, 102.9, 108.6–120.8 (4 C, m), 132.4, 135.5, 135.9, 136.1, 136.3, 136.7, 155.2. IR (CHCl₃): 1603, 1396, 1352, 1238, 1194, 1146 cm⁻¹. HRMS calcd for C₂₅H₂₈F₉O₃SSi₂ (M + H)⁺ *m/z*: 635.1154, found 635.1149.

■ ASSOCIATED CONTENT

● Supporting Information

Optimization of 3-borylbenzynes precursors (Tables S1–S3), interaction of ions with silyl- and borylbenzynes **4** and **5** (Tables S4–S6), ¹H and ¹³C NMR spectra for all new compounds, and Cartesian Coordinates for ground and transition state structures. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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Notes

The authors declare no competing financial interest.

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■ REFERENCES

- (1) (a) Eicher, T.; Hauptmann, S. *The Chemistry of Heterocycles*, Wiley-VCH: Weinheim, 2003. (b) Katrizky, A. R.; Pozharskii, A. F. *Handbook of Heterocyclic Chemistry*, 2nd ed., Pergamon: New York, 2000.
- (2) Saberi, M. R.; Vinh, T. K.; Yee, S. W.; Griffiths, B. J. N.; Evans, P. J.; Simons, C. J. *Med. Chem.* **2006**, *49*, 1016–1022.
- (3) Hayashi, S.; Hirao, A.; Imai, A.; Nakamura, H.; Murata, Y.; Ohashi, K.; Nakata, E. *J. Med. Chem.* **2009**, *52*, 610–625.
- (4) Patel, D.; Jain, M.; Shah, S. R.; Bahekar, R.; Jadav, P.; Darji, B.; Siriki, Y.; Bandyopadhyay, D.; Joharapurkar, A.; Kshirsagar, S.; Patel, H.; Shaikh, M.; Sairam, K. V. V. M.; Patel, P. *ChemMedChem* **2011**, *6*, 1011–1016.
- (5) Dixit, P. P.; Nair, P. S.; Patil, V. J.; Jain, S.; Arora, S. K.; Sinha, N. *Bioorg. Med. Chem. Lett.* **2005**, *15*, 3002–3005.
- (6) Rezaei, Z.; Khabnadideh, S.; Pakshir, K.; Hossaini, Z.; Amiri, F.; Assadpour, E. *Eur. J. Med. Chem.* **2009**, *44*, 3064–3067.
- (7) Dubey, A.; Srivastava, S. K.; Srivastava, S. D. *Bioorg. Med. Chem. Lett.* **2011**, *21*, 569–573.
- (8) Semple, G.; Skinner, P. J.; Cherrier, M. C.; Webb, P. J.; Sage, C. R.; Tamura, S. Y.; Chen, R.; Richman, J. G.; Connolly, D. T. *J. Med. Chem.* **2006**, *49*, 1227–1230.
- (9) Prous, J.; Graul, A.; Castañer, J. *Drugs Future* **1994**, *19*, 457–459.

(10) (a) Matsumoto, T.; Sohma, T.; Hatazaki, S.; Suzuki, K. *Synlett* **1993**, 843–846. (b) Jin, T.; Yamamoto, Y. *Angew. Chem., Int. Ed.* **2007**, *46*, 3323–3325. (c) Shi, F.; Waldo, J. P.; Chen, Y.; Larock, R. C. *Org. Lett.* **2008**, *10*, 2409–2412. (d) Liu, Z.; Shi, F.; Martinez, P. D. G.; Raminelli, C.; Larock, R. C. *J. Org. Chem.* **2008**, *73*, 219–226. (e) Dai, M.; Wang, Z.; Danishefsky, S. J. *Tetrahedron Lett.* **2008**, *49*, 6613–6616. (f) Campbell-Verduyn, L.; Elsinga, P. H.; Mirfeizi, L.; Dierckx, R. A.; Feringa, B. L. *Org. Biomol. Chem.* **2008**, *6*, 3461–3463. (g) Zhang, F.; Moses, J. E. *Org. Lett.* **2009**, *11*, 1587–1590. (h) Spiteri, C.; Keeling, S.; Moses, J. E. *Org. Lett.* **2010**, *12*, 3368–3371. (i) Dubrovskiy, A. V.; Larock, R. C. *Org. Lett.* **2010**, *12*, 1180–1183. (j) Li, P.; Zhao, J.; Wu, C.; Larock, R. C.; Shi, F. *Org. Lett.* **2011**, *13*, 3340–3343. (k) Fang, Y.; Wu, C.; Larock, R. C.; Shi, F. *J. Org. Chem.* **2011**, *76*, 8840–8851. (l) Lu, C.; Dubrovskiy, A. V.; Larock, R. C. *J. Org. Chem.* **2012**, *77*, 2279–2284.

(11) For recent reviews, see: (a) Pellissier, H.; Santelli, M. *Tetrahedron* **2003**, *59*, 701–730. (b) Wenk, H. H.; Winkler, M.; Sander, W. *Angew. Chem., Int. Ed.* **2003**, *42*, 502–528. (c) Dyke, A. M.; Hester, A. J.; Lloyd-Jones, G. C. *Synthesis* **2006**, *24*, 4093–4112. (d) Bhunia, A.; Yetra, S. R.; Biju, A. T. *Chem. Soc. Rev.* **2012**, *41*, 3140–3152. (e) Tadross, P. M.; Stoltz, B. M. *Chem. Rev.* **2012**, *112*, 3550–3577.

(12) Garg and Houk et al. and our group have independently studied the (3 + 2) cycloaddition reactions of 3-silylbenzynes with 1,3-dipoles along with computational analyses, and both of our groups have come to essentially the same conclusions. See: Bronner, S. M.; Mackey, J. L.; Houk, K. N.; Garg, N. K. *J. Am. Chem. Soc.* **2012**, *134*, 13966–13969.

(13) Very recently, Garg et al. achieved regioselective (3 + 2) cycloadditions of pyridynes possessing either a sulfamoyloxy group or a bromo group as a directing group; see: Goetz, A. E.; Garg, N. K. *Nat. Chem.* **2013**, *5*, 54–60.

(14) (a) Akai, S.; Ikawa, T.; Takayanagi, S.; Morikawa, Y.; Mohri, S.; Tsubakiyama, M.; Egi, M.; Wada, Y.; Kita, Y. *Angew. Chem., Int. Ed.* **2008**, *47*, 7673–7676. (b) Ikawa, T.; Takagi, A.; Kurita, Y.; Saito, K.; Azechi, K.; Egi, M.; Kakiguchi, K.; Kita, Y.; Akai, S. *Angew. Chem., Int. Ed.* **2010**, *49*, 5563–5566. (c) Ikawa, T.; Nishiyama, T.; Shigeta, T.; Mohri, S.; Morita, S.; Takayanagi, S.; Terauchi, Y.; Morikawa, Y.; Takagi, A.; Ishikawa, Y.; Fujii, S.; Kita, Y.; Akai, S. *Angew. Chem., Int. Ed.* **2011**, *50*, 5674–5677.

(15) A part of this work was reported in our very recent brief account; see: Ikawa, T.; Tokiwa, H.; Akai, S. *J. Synth. Org. Chem. Jpn.* **2012**, *70*, 1123–1133.

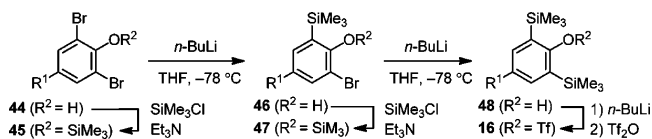
(16) Pansegrau, P. D.; Rieker, W. F.; Meyers, A. I. *J. Am. Chem. Soc.* **1988**, *110*, 7178–7184.

(17) (a) Sapountzis, I.; Lin, W.; Fischer, M.; Knochel, P. *Angew. Chem., Int. Ed.* **2004**, *43*, 4364–4366. (b) Lin, W.; Chen, L.; Knochel, P. *Tetrahedron* **2007**, *63*, 2787–2797.

(18) 2,4,5-Trichlorobenzenesulfonyl chloride, used for preparing **14B**, is readily available from several chemical companies, e.g., TCI Fine Chemicals, TCI America, Acros Organics, Sigma-Aldrich, and Alfa Aesar.

(19) According to our previous work,^{14b} new borylbenzyne precursors **14** were prepared from 2,6-diiodophenols **28**. For details, see experimental section and reference 14b.

(20) According to our previous work,^{14c} silylbenzyne precursors **16** were prepared from 2,6-dibromophenols **44** as shown in the scheme below. For details, see reference 14c.



(21) Allred, A. L.; Rochow, E. G. *J. Inorg. Nucl. Chem.* **1958**, *5*, 264–268.

(22) Chattopadhyay, S.; Chattopadhyay, U.; Mathur, P. P.; Saini, K. S.; Ghosal, S. *Planta Med.* **1983**, *49*, 252–254.

(23) Ikawa, T.; Nishiyama, T.; Nosaki, T.; Takagi, A.; Akai, S. *Org. Lett.* **2011**, *13*, 1730–1733.

(24) Dhanak, D.; Knight, S. D. Patent WO 2007/103755 A2.

(25) Kim, S. W.; Biegion, A.; Katsamanis, Z. E.; Ehrlich, C. W.; Hooker, J. M.; Shea, C.; Muench, L.; Xu, Y.; King, P.; Carter, P.; Alexoff, D. L.; Fowler, J. S. *Nucl. Med. Biol.* **2009**, *36*, 323–334.

(26) These deborylation conditions (soon to be published) were recently found in our laboratory.

(27) All DFT calculations were performed at the B3LYP/6-31G(d) level except for iodine atom (B3LYP/LanL2DZ) of **6k**, **TS17**, and **TS18**. Optimized structures of the reactants, transition states, and products were characterised by analytical frequency calculations, and all the total electronic energies were included in the zero-point energy corrections at the same level. The calculated number of imaginary frequencies (NImag) will determine whether the optimized structures are the energy minima (NImag = 0) or transition states (NImag = 1) along the reaction pathway. All calculations were carried out by using the Gaussian 09 revision A.02 (Frisch, M. J. et al.; see the Supporting Information for full citation).

(28) The interaction of ions in the reaction mixture with benzynes (**4** and **5**) may have some influence on the regioselectivities. However, in fact, the ratios of distal- to proximal-adducts were not significantly affected by changing the conditions of preparation of **4** and **5** (see Tables S5 and S6 in Supporting Information). Therefore, we have reached the conclusion that the effects of ions were negligibly small, and they were not considered for the calculations of ground and transition state structures.

(29) (a) Cheong, P. H.-Y.; Paton, R. S.; Bronner, S. M.; Im, G.-Y. J.; Garg, N. K.; Houk, K. N. *J. Am. Chem. Soc.* **2010**, *132*, 1267–1269.

(b) Im, G.-Y.; Bronner, S. M.; Goetz, A. E.; Paton, R. S.; Cheong, P. H.-Y.; Houk, K. N.; Garg, N. K. *J. Am. Chem. Soc.* **2010**, *132*, 17933–17944.

(30) In general, accurate entropy and thermodynamic corrections on theoretical reaction analysis should be required (see: Kruse, H.; Goerigk, L.; Grimme, S. *J. Org. Chem.* **2012**, *77*, 10824–10834. Grimme, S. *ChemPhysChem* **2012**, *13*, 1407–1409). However, $\Delta\Delta H^\ddagger$ values of the (3 + 2) cycloaddition reactions of borylbenzynes **4** without entropy contributions were applied for calculating the theoretical ratios of distal- to proximal-adducts because it is difficult to correctly evaluate the activation entropy for the reactions of **4** due to local interactions between the boryl group and solvent molecules. On the other hand, $\Delta\Delta G^\ddagger$ values of the (3 + 2) cycloaddition reactions of silylbenzynes **5** were applied for the calculation.

(31) As for the (3 + 2) cycloaddition of 3-silylbenzynes, similar discussion was also reported by Houk and Garg.¹²

(32) All graphics were prepared with CYLview: CYLview, 1.0b; Legault, C. Y. Université de Sherbrooke, 2009 (<http://www.cylview.org>).

(33) (a) Miyaura, N.; Suzuki, A. *Chem. Rev.* **1995**, *95*, 2457–2483. (b) Monnier, F.; Taillefer, M. *Angew. Chem., Int. Ed.* **2009**, *48*, 6954–6971.

(34) Maddani, M. R.; Moorthy, S. K.; Prabhu, K. R. *Tetrahedron* **2010**, *66*, 329–333.

(35) Hubbard, A.; Okazaki, T.; Laali, K. K. *J. Org. Chem.* **2007**, *73*, 316–319.

(36) Liu, C.-Y.; Knochel, P. *J. Org. Chem.* **2007**, *72*, 7106–7115.

(37) Nguyen, T.-T.-T.; Simon, F.-X.; Schmutz, M.; Mésini, P. *J. Chem. Commun.* **2009**, 3457–3459.

(38) Murali, A.; Puppala, M.; Varghese, B.; Baskaran, S. *Eur. J. Org. Chem.* **2011**, 5297–5302.

(39) Lamani, M.; Prabhu, K. R. *Angew. Chem., Int. Ed.* **2010**, *49*, 6622–6625.

(40) Young, M. B.; Barrow, J. C.; Glass, K. L.; Lundell, G. F.; Newton, C. L.; Pellicore, J. M.; Rittle, K. E.; Selnick, H. G.; Stauffer, K. J.; Vacca, J. P.; Williams, P. D.; Bohn, D.; Clayton, F. C.; Cook, J. J.; Krueger, J. A.; Kuo, L. C.; Lewis, S. D.; Lucas, B. J.; McMasters, D. R.; Miller-Stein, C.; Pietrak, B. L.; Wallace, A. A.; White, R. B.; Wong, B.; Yan, Y.; Nantermet, P. G. *J. Med. Chem.* **2004**, *47*, 2995–3008.

(41) Javed, M. I.; Brewer, M. *Org. Synth.* **2008**, *85*, 189–195.

(42) Katritzky, A. R.; Button, M. A. C.; Denisenko, S. N. *J. Heterocycl. Chem.* **2000**, *37*, 1505–1510.

- (43) Cossy, J.; Tresnard, L.; Pardo, D. G. *Eur. J. Org. Chem.* **1999**, 1925–1933.
- (44) Still, W. C.; Kahn, M.; Mitra, A. *J. Org. Chem.* **1978**, *43*, 2923–2925.
- (45) Clary, J. W.; Rettenmaier, T. J.; Snelling, R.; Bryks, W.; Banwell, J.; Wipke, W. T.; Singaram, B. *J. Org. Chem.* **2011**, *76*, 9602–9610.